



## DAFTAR PUSTAKA

- Al-Mamoori, A., Krishnamurthy, A., Rownaghi, A. A., & Rezaei, F. (2017). Carbon Capture and Utilization Update. In *Energy Technology* (Vol. 5, Issue 6, pp. 834–849). Wiley-VCH Verlag. <https://doi.org/10.1002/ente.201600747>
- Azzahra, Aulia. (2025). Studi Kinetika Sekuestrasi Karbon Dioksida Menggunakan Red Mud. Tesis. Departemen Teknik Kimia Fakultas Teknik Universitas Gadjah Mada
- Bagusantara, Fathin. (2024). Studi Kinetika Penangkapan CO<sub>2</sub> menggunakan *Red Mud* dengan Variasi *Solid to Liquid Ratio Red Mud* dan Komposisi Gas CO<sub>2</sub>. Laporan Penelitian. Departemen Teknik Kimia Fakultas Teknik Universitas Gadjah Mada
- Bonenfant, D., Kharoune, L., Sauv e, S., Hausler, R., Niquette, P., Mimeault, M., & Kharoune, M. (2008). CO<sub>2</sub> sequestration by aqueous red mud carbonation at ambient pressure and temperature. *Industrial and Engineering Chemistry Research*, 47(20), 7617–7622. <https://doi.org/10.1021/ie7017228>
- Calvin, K., Dasgupta, D., Krinner, G., Mukherji, A., Thorne, P. W., Trisos, C., Romero, J., Aldunce, P., Barrett, K., Blanco, G., Cheung, W. W. L., Connors, S., Denton, F., Diongue-Niang, A., Dodman, D., Garschagen, M., Geden, O., Hayward, B., Jones, C., ... Ha, M. (2023). *IPCC, 2023: Climate Change 2023: Synthesis Report. Contribution of Working Groups I, II and III to the Sixth Assessment Report of the Intergovernmental Panel on Climate Change [Core Writing Team, H. Lee and J. Romero (eds.)]. IPCC, Geneva, Switzerland.* (P. Arias, M. Bustamante, I. Elgizouli, G. Flato, M. Howden, C. M endez-Vallejo, J. J. Pereira, R. Pichs-Madruga, S. K. Rose, Y. Saheb, R. S anchez Rodr iguez, D.  urge-Vorsatz, C. Xiao, N. Yassaa, J. Romero, J. Kim, E. F. Haites, Y. Jung, R. Stavins, ... C. P ean, Eds.). <https://doi.org/10.59327/IPCC/AR6-9789291691647>
- Darmana, D., Henket, R. L. B., Deen, N. G., & Kuipers, J. A. M. (2007). Detailed modelling of hydrodynamics, mass transfer and chemical reactions in a bubble column using a discrete bubble model: Chemisorption of CO<sub>2</sub> into NaOH solution, numerical and experimental study. *Chemical Engineering Science*, 62(9), 2556–2575. <https://doi.org/10.1016/j.ces.2007.01.065>
- Han, Y. S., Ji, S., Lee, P. K., & Oh, C. (2017). Bauxite residue neutralization with simultaneous mineral carbonation using atmospheric CO<sub>2</sub>. *Journal of Hazardous Materials*, 326, 87–93. <https://doi.org/10.1016/j.jhazmat.2016.12.020>
- Hemmati, A., Shayegan, J., Bu, J., Yeo, T. Y., & Sharratt, P. (2014). Process optimization for mineral carbonation in aqueous phase. *International Journal of Mineral Processing*, 130, 20–27. <https://doi.org/10.1016/j.minpro.2014.05.007>
- Johnston, M., Clark, M. W., McMahon, P., & Ward, N. (2010). Alkalinity conversion of bauxite refinery residues by neutralization. *Journal of Hazardous Materials*, 182(1–3), 710–715. <https://doi.org/10.1016/j.jhazmat.2010.06.091>



- Kaithwas, A., Prasad, M., Kulshreshtha, A., & Verma, S. (2012). Industrial wastes derived solid adsorbents for CO<sub>2</sub> capture: A mini review. *Chemical Engineering Research and Design*, 90(10), 1632–1641. <https://doi.org/10.1016/j.cherd.2012.02.011>
- Kashefi, K., Pardakhti, A., Shafiepour, M., & Hemmati, A. (2020). Process optimization for integrated mineralization of carbon dioxide and metal recovery of red mud. *Journal of Environmental Chemical Engineering*, 8(2). <https://doi.org/10.1016/j.jece.2019.103638>
- Krauß, M., & Rzehak, R. (2017). Reactive absorption of CO<sub>2</sub> in NaOH: Detailed study of enhancement factor models. *Chemical Engineering Science*, 166, 193–209. <https://doi.org/10.1016/j.ces.2017.03.029>
- Li, L., Liu, Q., Huang, T., & Peng, W. (2022). Mineralization and utilization of CO<sub>2</sub> in construction and demolition wastes recycling for building materials: A systematic review of recycled concrete aggregate and recycled hardened cement powder. *Separation and Purification Technology*, 298. <https://doi.org/10.1016/j.seppur.2022.121512>
- Liu, S., Shen, Y., Wang, Y., Shen, P., Xuan, D., Guan, X., & Shi, C. (2022). Upcycling sintering red mud waste for novel superfine composite mineral admixture and CO<sub>2</sub> sequestration. *Cement and Concrete Composites*, 129. <https://doi.org/10.1016/j.cemconcomp.2022.104497>
- Lyu, F., Hu, Y., Wang, L., & Sun, W. (2021). Dealkalization processes of bauxite residue: A comprehensive review. *Journal of Hazardous Materials*, 403(May 2020), 123671. <https://doi.org/10.1016/j.jhazmat.2020.123671>
- Mitchell, M. J., Jensen, O. E., Cliffe, K. A., & Maroto-Valer, M. M. (2010). A model of carbon dioxide dissolution and mineral carbonation kinetics. *Proceedings of the Royal Society A: Mathematical, Physical and Engineering Sciences*, 466(2117), 1265–1290. <https://doi.org/10.1098/rspa.2009.0349>
- Pohorecki, R., & Moniuk, W. (1988). Kinetics of reaction between carbon dioxide and hydroxyl ions in aqueous electrolyte solutions. *Chemical Engineering Science*, 43(7), 1677–1684. [https://doi.org/10.1016/0009-2509\(88\)85159-5](https://doi.org/10.1016/0009-2509(88)85159-5)
- Qin, J., Ying, J., Wang, Y., Niu, A., Lin, C., Qiu, R., & Lim, J. W. (2023). Insights into active and passive carbon sequestration and causticity reduction in hazardous red mud slurry. *Carbon Research*, 2(1). <https://doi.org/10.1007/s44246-023-00071-3>
- Rahmanihazaki, M., & Hemmati, A. (2022). A review of mineral carbonation by alkaline solidwaste. In *International Journal of Greenhouse Gas Control* (Vol. 121). Elsevier Ltd. <https://doi.org/10.1016/j.ijggc.2022.103798>
- Sahu, R. C., Patel, R. K., & Ray, B. C. (2010). Neutralization of red mud using CO<sub>2</sub> sequestration cycle. *Journal of Hazardous Materials*, 179(1–3), 28–34. <https://doi.org/10.1016/j.jhazmat.2010.02.052>
- Su, Z., Wang, D., Huang, X., Li, X., Yu, X., Huang, Y., & Feng, Q. (2020). Mechanism of CO<sub>2</sub> sequestration by high-alkaline red mud: Conversion of katoite. *International Journal of Global Warming*, 22(2), 160–173. <https://doi.org/10.1504/IJGW.2020.110293>



- Wang, X., Zhan, Q., Zhang, X., Su, Y., & Zhou, J. (2024). Study on improving the carbon sequestration properties of sintered red mud by regulating pressure. *Journal of Building Engineering*, 84. <https://doi.org/10.1016/j.jobbe.2024.108518>
- Yadav, V. S., Prasad, M., Khan, J., Amritphale, S. S., Singh, M., & Raju, C. B. (2010). Sequestration of carbon dioxide (CO<sub>2</sub>) using red mud. *Journal of Hazardous Materials*, 176(1–3), 1044–1050. <https://doi.org/10.1016/j.jhazmat.2009.11.146>
- Yang, J., Xiao, H., He, X., Su, Y., Zeng, J., Li, W., Li, Y., & Qi, H. (2024). Rapid wet grinding carbonation of sintering red mud for highly efficient CO<sub>2</sub> sequestration and Cr solidification. *Chemical Engineering Journal*, 488. <https://doi.org/10.1016/j.cej.2024.151134>
- Yoo, M., Han, S. J., & Wee, J. H. (2013). Carbon dioxide capture capacity of sodium hydroxide aqueous solution. *Journal of Environmental Management*, 114, 512–519. <https://doi.org/10.1016/j.jenvman.2012.10.061>
- Zeebe, Richard E., Dieter Wolf-Gladrow. (2001). CO<sub>2</sub> in Seawater: Equilibrium, Kinetics, Isotopes. Elsevier Oceanography Series.
- Zhang, X., Wang, X., Zhan, Q., & Song, C. (2024). Carbonation reaction properties and reaction mechanisms of red mud under different carbon dioxide pressures. *Journal of Environmental Chemical Engineering*, 12(3). <https://doi.org/10.1016/j.jece.2024.112910>